Interpreting Balanced Chemical Equations name:

Directions: For the following chemical equations balance each and then write out ALL mole ratios.

1. K(s) + H2O(l) 🡪 KOH(aq) + H2(g)
2. C2H5OH(l) + O2 🡪 CO2(g) + H2O(l)
3. Bi(NO3)(aq) + H2S(aq) 🡪 Bi2S3(s) + HNO3(aq)
4. Al(s) + Br2(l) 🡪 AlBr3(s)

Directions: For the following questions DO NOT SOLVE, just write out the conversion factors you would need to use to solve for the following balanced chemical equation.

2Rb(s) + H2O(l) 🡪 H2(g) +2RbOH(aq)

1. If there is 2.0g of solid Rubidium, how many mols of water are needed?
2. If you have 3.56 mols of solid Rubidium, how many mols of Rubidium hydroxide would be produced?
3. If you have 4.2 grams of solid Rubidium, how any grams of Hydrogen gas will be produced?