Atomic Review WS name:

1. Label the parts of the atom below

 

1. Fill in the table below

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Atom | Atomic # | Mass # | Neutrons |  Protons | Electrons |
| Aluminum |  |  |  |  |  |
| BariumXenon |  |  |  |  |  |
| Tungsten |  |  |  |  |  |

1. Fill in the table below

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Isotope | Atomic # | Mass # | Neutrons |  Protons | Electrons |
| germanium-64 |  |  |  |  |  |
| neon-23 |  |  |  |  |  |
| carbon-12carbon-14 |  |  |  |  |  |

1. The element of copper has naturally occurring isotopes with mass numbers of 64 and 65. The relative abundance and atomic masses are 69.2% for mass = 62.93 amu, and 30.8% for mass = 64.93 amu. Calculate the average atomic mass of copper.