**Practice Problems on Net Ionic Equations:**

Show the total ionic and net ionic forms of the following equations.  If all species are spectator ions, please indicate that no reaction takes place.  Note!  You need to make sure the original equation is balanced before proceeding!

1.  AgNO3 (aq)   +   KCl (aq)         AgCl (s)   +   KNO3 (aq)

2.  Mg(NO3)2 (aq)    +   Na2CO3 (aq)       MgCO3 (s)   +     NaNO3 (aq)

3.  strontium bromide (aq)  +  potassium sulfate (aq)       strontium sulfate (s)   +   potassium bromide (aq)

4.  manganese(II) chloride (aq)   +    ammonium carbonate (aq)      manganese(II) carbonate (s)  +   ammonium chloride (aq)

5.  chromium(III) nitrate (aq)    +   iron(II) sulfate (aq)     chromium(III) sulfate (aq)   +  iron(II) nitrate (aq)

Please complete the following reactions, and show the total ionic and net ionic forms of the equation by predicting the products. Use your solubility rules to do the problems correctly.

6.  K3PO4 (aq)   +   Al(NO3)3 (aq)   

7.  BeI2 (aq)   +   Cu2SO4 (aq)   

8.  Ni(NO3)3 (aq)    +   KBr (aq)   

9.  cobalt(III) bromide   +   potassium sulfide    

10.  barium nitrate    +     ammonium phosphate    