Titration Note Taking Guide

1) What is the learning target?

2) Example: A 25mL solutions of .5M NaOH is titrated until neutralized into a 50 mL sample of HCl. What was the concentration of the HCl?

M(acid)X V(acid)= M(base)X V(acid)

Titration WSQ

Watch: Key Concepts:

What is the purpose of a titration?

What equipment is needed for a titration?

What you should you do with your burette first?

What do you do if there are air bubbles on the burette?

Should you need to read the initial volume of the burette?

How many drops of indicator do you add?

To get the estimated volume of Titrant needed what do you need to do?

What color should your solution be if you performed the titration perfectly?

Summary:

Question: